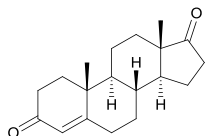


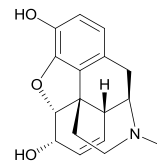
Strychnine  
(poison)



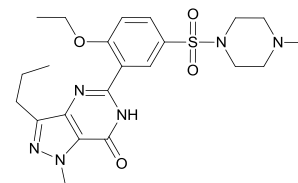
Androstenedione  
(estrogen and testosterone  
precursor)

## Organic Chemistry Chemistry 143

Fall 2015



Morphine  
(opiate analgesic  
drug)



Viagra  
(vasodilation by  
inhibition of PDE5)

**Instructor:** Matthias Brewer 6-1042 Cook A316 Matthias.Brewer@uvm.edu

**Lecture:** 8:30am – 9:20am MWF, Jeffords Hall, Rm 127

**Laboratory:** Thursday 8:30 – 11:15 AM and 1:15 – 4:00 PM

**ACS Placement Exam will be given in Laboratory on Sept. 3<sup>rd</sup>**

**Laboratory Check-in / ChemDraw Tutorial – Sept. 10<sup>th</sup>**

### Required Course Materials:

Organic Chemistry 6th ed., Loudon and Parise, Roberts and Co., 2016 (ISBN: 9781936221639)

Organic Chemistry Study Guide and Solutions Manual 6<sup>th</sup> ed., Loudon and Parise, Roberts and Co., 2016

Sapling Learning account: account can be purchased as part of book order.

### Recommended Course Materials:

Molecular Structure Models (e.g.: ISBN: 0471-362719)

### Books on reserve in library that may be useful:

Organic Chemistry I as a Second Language: Translating the Basic Concepts 2<sup>nd</sup> ed., D. Klein; ISBN: (978-0470-12929-6)

Organic Chemistry II as a Second Language: Second Semester Topics 2<sup>nd</sup> ed., D. Klein; ISBN: (978-0-471-73808-4)

The Art of Writing Reasonable Organic Reaction Mechanisms R.B. Grossman ISBN:0-387-95468-6

Writing Reaction Mechanisms in Organic Chemistry A. Miller ISBN: 0-12-496711-6

**Course Prerequisite:** Chemistry 32 or 36.

### Office hours:

Monday 11:30 – 12:30

Wednesday 11:30 – 12:30

Thursday 11:30 – 12:30

or by appointment

### This course will address learning goals 1, 2, 3, and 5 below for chemistry majors:

1. Students will demonstrate general knowledge in chemistry and will be able to apply chemical and physical principles in the solution of qualitative and quantitative chemical problems.
2. Students will understand the interplay of observational data, hypotheses, and hypothesis-driven experimentation through application of the scientific method.
3. Students will become proficient in chemical laboratory techniques and be able to apply these to practical and current problems in research.
4. Students will be able to read and critically evaluate the chemical and scientific literature.

5. The students will learn to present scientific data clearly and effectively through both written and verbal communication.

### General Comments

In Chemistry 143 we begin an exploration of the basic principles of Organic Chemistry. You will find that Organic Chemistry involves many new concepts, a large number of rules and (by the end of the second semester) a large number of reaction mechanisms. However, as the course progresses and your knowledge grows, you will find that a relatively small subset of concepts tie together the vast amount of information contained in the text. Learning these underlying principles, and knowing when and how to apply them to solve problems, is the key to success. You have seen many of these concepts in General Chemistry, but here they will be considered from a different point of view. For example, knowing the relative electronegativity of atoms is essential to understanding why molecules react the way they do; the concept of electronegativity allows you to rationalize why some atoms are good leaving groups and others are not. *A special effort made at the beginning of the course to master the writing of proper Lewis structures with the correct number of bonds, formal charges, and unshared pairs of electrons will pay off as the course progresses. Also, an early and thorough understanding of the relative electronegativity of atoms, Lewis acid-base theory, Bronstead-Lowry acid-base theory, and the rules for writing proper contributing "structures" to resonance hybrids will make the understanding of reaction mechanisms considerably easier.*

For each chapter you should work as many of the suggested problems as possible and I strongly urge you to keep up with your reading and problem solving. Organic Chemistry is not inherently difficult, but it is different than any chemistry you have seen thus far. You will need to understand new concepts and then apply those concepts in new situations, which will take a slow and steady approach...*cramming does not work well in this subject!*

### Academic Conduct:

Cheating will be considered grounds for failing the course. All graded assignments must be your own work. **This includes on-line homework assignments which are independent exercises, not group exercises.** Cases of cheating or plagiarism *will* lead to further disciplinary action which may include dismissal from the University according to the rules set forth in The University of Vermont's *Code of Academic Integrity*.

### Grading:

Your course grade will be based on ten on-line homework assignments, three examinations, a cumulative final examination, and your laboratory grade. (**Note:** You must earn a passing grade in the laboratory to receive a passing grade for the course. More than two laboratories missed for any reason will result in a failing grade for the course unless you are granted an incomplete by your Dean).

Lab	20%
Homework	15%
3 Midterm Exams	45%
Cumulative Final	20%

**Midterm Dates:**

Thursday, October 1	6:00 P.M.-8:00 P.M.	Kalkin 004
Thursday, October 22	6:00 P.M.-8:00 P.M.	Kalkin 004
Thursday, November 19	6:00 P.M.-8:00 P.M.	Kalkin 004

**Final Exam Date:**

Friday, Dec 18th	10:30 A.M.-1:15 P.M.	Jeffords 127
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**On-line homework is due by 11:59 pm on the following dates:**

1. September 6
2. September 13
3. September 20
4. September 27
5. October 11
6. October 18
7. November 1
8. November 8
9. November 15
10. December 6

No exam grades are dropped. The only valid excuses for missing an exam are medical or other true emergency situations. If you miss an exam for such a reason, you must inform me of it promptly, present appropriate documentation of your excuse, and receive formal approval to take a make up exam. If you miss an exam for any other reason, you will receive a grade of zero for that exam. The answers to exam problems will be posted after each exam. If you have any questions concerning the grading of an exam, you must see me within one week after the day the exam is returned to the class. Exams must be taken in ink to insure that you can get points for a grading error.

The lowest quiz score will be dropped and will be replaced by the average score of the nine remaining quizzes.

*Religious Holidays: Students have the right to practice the religion of their choice. Each semester students should submit in writing to their instructors by the end of the second full week of classes their documented religious holiday schedule for the semester. Faculty must permit students who miss work for the purpose of religious observance to make up this work.*

**Tentative Outline of Course**

Chapter 1. Chemical Bonding and Chemical Structure.

Sections: All sections

Suggested Problems:

Chapter 3. Acids and Bases: The curved arrow notation

Sections: 3.1-3.6

Suggested Problems:

Chapter 2. Alkanes.

Sections: 2.1-2.5, 2.8

Suggested Problems:

Chapter 4. Introduction to Alkenes: Structure and Reactivity

Sections: All sections

Suggested Problems:

Chapter 5. Addition Reactions of Alkenes

Sections: All sections

Suggested Problems:

Chapter 14. The Chemistry of Alkynes

Sections: 14.1, 14.2, 14.4-14.7

Suggested Problems:

Chapter 6. Principles of Stereochemistry

Sections: All sections

Suggested Problems:

Chapter 7. Cyclic Compounds: Stereochemistry of reactions

Sections: All Sections

Suggested Problems:

Chapter 8. Noncovalent Intermolecular Interactions

Sections: 8.1-8.3

Suggested Problems:

Chapter 9. The Chemistry of Alkyl Halides

Sections: All Sections

Suggested Problems:

Chapter 10. The Chemistry of Alcohols and Thiols

Sections: 10.1-10.7

Suggested Problems:

## Chapter 11. The Chemistry of Ethers, Epoxides, Glycols, and Sulfides.

Sections 11.1-11.6, 11.8, 11.10, and Chapter 14 Section 14.8

Suggested Problems:

Concepts you must understand from General Chemistry:

- Properties of covalent bonds
- The octet rule
- Structural isomers
- Lewis dot structures
- Formal charges
- Resonance
- Electronegativity and bond polarity
- VSEPR (Valence Shell Electron Pair Repulsion)
- Hybridization

Key's to success in Organic Chemistry:

- Do not try to cram!
- You will see many new concepts in this course. Try to write out an explanation of the concepts in your own words as if explaining them to someone else.
- Work as many practice problems as possible. Practice problem reinforce the new concepts and are the only way to test your understanding of the material.
- Do not look at a problem's answer until you have really tried the problem. After seeing the answer it often seems obvious and you may assume you understand.
- When you get a problem wrong, try to understand where your thinking was in error and attempt to identify what concept you missed.
- Ask questions!
- Come to office hours or make an appointment with me to resolve any questions early!
- Review the material frequently.