CHEM 36 General Chemistry Quiz #5

March 8, 2002

Name <u>:</u>	
	The pH of a solution is 5.00 calculate the concentration (in mol/L) of H ⁺ in the solution.
2.	The pOH of a solution is 5.00 calculate the concentration (in mol/L) of H^{+} in the solution.
3.	The conjugate acid of NH $_3$ is the ammonium ion (NH $_4$ ⁺). If K $_b$ for ammonia is equal to 1.8 x 10 $^{-5}$, is the ammonium ion a stronger or weaker acid than acetic acid (K $_a$ = 1.8 x 10 $^{-5}$)? Explain.